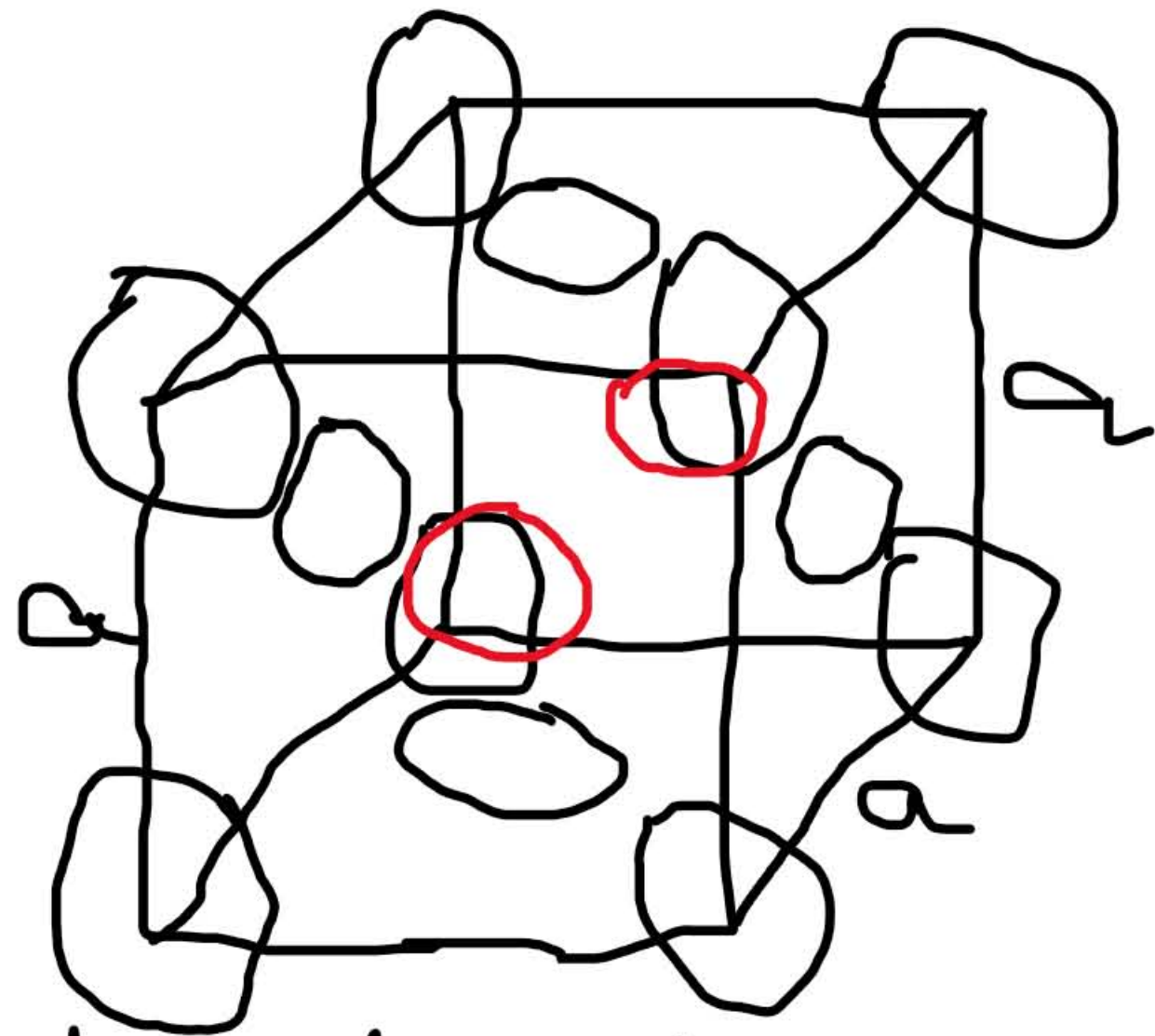




$$\rho = 1 \text{ g/cm}^3$$

$$\rho = \frac{m}{V} = 8,9 \text{ g/cm}^3$$

$$6 \times \frac{1}{2} + 8 \times \frac{1}{8} = 4 \text{ Átomos}$$

Cu CABRE CFC

$$a = 0,128 \text{ nm} = 0,128 \times 10^{-7} \text{ cm}$$

$$M = 63,55 \text{ g/mol}$$

$$1 \text{ mol} = 6,02 \times 10^{23} \text{ Átomos}$$

$$M = 63,55 \text{ g} / 6,02 \times 10^{23} \text{ Átomos}$$

$$\rho = \frac{m}{V} = \frac{4 \times 63,55}{(0,128 \times 10^{-7} \text{ cm})^3} = 8,9 \text{ g/cm}^3$$